

# Copper-diimine coordination compounds as potential new tools in the treatment of cancer

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## Abstract

The usefulness of coordination metal compounds in cancer chemotherapy was firstly demonstrated by Cisplatin. As a result of an intense research on the field of copper complexes with antitumor activity, copper coordination compounds are emerging as an alternative for the development of drugs for the treatment cancer. Our research is devoted to find new metal containing drugs for the treatment of cancer. This paper summarizes our results on the study of Copper-diimine of ternary complexes and related Copper homoleptic complexes as cytotoxic agents. The studied diimines were 2,2-dipyridil-amine, bipyridine, 4,4'-2,2 bipyridine, 1,10 phenanthroline, 4-methyl-1,10 phenanthroline, 5-NO<sub>2</sub>-phenanthroline, neocuproine and bathophenanthroline. L-dipeptides, iminodiacetic acid or phosphines were used as ligands. Most of the complexes present higher cytotoxicity than Cisplatin. We hypothesize that the cation [Cu(diimine)]<sup>2+</sup> is the active species of the complexes.

## Introduction

The usefulness of coordination metal compounds in cancer chemotherapy was firstly demonstrated by Cisplatin, *cis*-diamminedichloridoplatinum(II) [1]. While the anticancer activity of this drug was somewhat serendipitously discovered in the '60s still is a widely used chemotherapeutic drug to treat cancers. Its clinical applications are limited by its side effects and drug resistance. The discovery of the activity of Cisplatin started an extensive research on Pt compounds that yielded several other compounds with antitumor activity, some of them in clinical use as Carboplatin and Oxaliplatin [2-4].

Complexes of other metals than Pt started to be studied, many with promising results. Different metal complexes may present different mechanism of action than Cisplatin and congeners and therefore different spectra of activity [3-6].

In this search for complexes with antitumor activity one strategy was to study coordination compounds of essential metals. This research was initially performed considering that living organisms already have metabolic pathways for these metals, and therefore may present less severe side effects. Different Cu(II) and Cu(I) complexes with antitumor activity have been synthesized and characterized as reviewed [7-9], some with encouraging results. Many Cu (II) complexes are active in spite of its ligands not having appreciable cytotoxic activity. As a result, copper coordination compounds are emerging as an alternative for the development of drugs for the treatment of cancer.

To the date, two copper compounds are about to start clinical trials. Casiopeina III-ia [(Cu(II))(4,4'-dimethyl-2,2'-bipyridine)(acetylacetonate)(NO<sub>2</sub>)(H<sub>2</sub>O)], developed by L. Ruiz and co-workers has completed the preclinical trials required to start clinical phase I studies in Mexico [10,11]. HydroCuP<sup>+</sup>, the phosphino copper(I) complex [Cu(tris-hydroxymethylphosphine)<sub>4</sub>][PF<sub>6</sub>], presents highly selective antitumor effects, showing promising results in latter preclinical studies [12,13].

The mechanism of action of copper compounds is not completely understood and possibly includes different molecular events [7,14,15]. Copper presents two oxidation states +I and +II which gives rise to a rich redox chemistry that may lead to the production of reactive oxygen species (ROS). ROS production is considered to be a relatively selective mechanism, as cancer cells are more vulnerable to oxidative stress than normal cells [16]. There are studies that relate cytotoxic activity with an increased intracellular copper uptake induced by the complexes [17,18]. According to those observations, Cu(I/II) would be the final active species, being the complex a "carrier" of Cu to the cell. The proposed mechanism of action of Casiopeinas<sup>®</sup> includes DNA binding. ROS production and oxidative damage of relevant biomolecules, leading to apoptosis of the cells [11,19]. HydroCuP<sup>+</sup> induces parapototic cell death due to endoplasmatic reticulum stress [13].

In relation to the diiminc ligands, 1,10-phenanthroline (phen) is one of the more common ligands used in coordination chemistry as it forms stable complexes with several metallic cations. Homoleptic Cu-complexes, [Cu(phen)<sub>2</sub>]<sup>+</sup> and [Cu(phen)<sub>2</sub>]<sup>2+</sup> bind to the minor groove of the double stranded DNA noncovalently, cleaving DNA at its binding region [20-23]. In addition, a number of heteroleptic phen-containing Cu (II) complexes show DNA binding and nuclease activity. Many of them present cytotoxic activity on tumor cell lines [24,25]. The 5-NO<sub>2</sub>-1,10-phenanthroline (5-NO<sub>2</sub>-phen) ligand presents different hydrophobicity and hydrogen bonding properties than the 1,10-phenanthroline (phen). It acts as an electron-tractor substituent, modifying the electronic properties of the molecule. These properties may influence their complexes reactivity, possibly showing

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different ability to bind to DNA than Cu-phen complexes [20-23]. As a result, their cytotoxic activity may be different. This difference possibly extends to heteroleptic 5-NO<sub>2</sub>-phen complexes.

Our research is devoted to find new metal containing compounds which may be useful for the treatment of cancer. In this paper we summarize our results on the study of Copper-diimine heteroleptic complexes and related Copper homoleptic complexes as cytotoxic agents. Figure 1 presents the diiminic ligands studied in this work. As anionic ligands, in order to form neutral complexes, L-dipeptide, diiminodiacetate (ida) were used. For the Cu(I) complexes phosphines were used as coligands.

### Cytotoxic activity of Copper complexes containing diimines

Table 1 [28-30] summarizes the results we have already obtained for different studied Cu-diimine containing complexes. Most of them present increased cytotoxic activity when compared to Cisplatin. Results are analyzed according to compound chemical groups.

#### Homoleptic CuL-dipeptide and Cu-diimine complexes

We started our studies on cytotoxic copper compounds preparing Cu-L-dipeptide complexes.

In relation to their possible mechanism of action, we studied their binding to (isolated) Calf Thymus DNA. EPR studies showed that the

**Table 1.** Cytotoxic activity expressed by IC<sub>50</sub> (unless specified) of the studied complexes against MDA-MB-23 (human metastatic breast cancer), MCF-7 (human metastatic breast adenocarcinoma) HeLa (human cervical adenocarcinoma), and A549 (human lung epithelial carcinoma) cell lines

Compound	IC50/μM MCF-7	IC50/μM MDA-MB-231	IC50/μM HeLa	IC50/μM A549	Ref.
Cu(L-Ala-Phe)	100	200	-	-	[26]
Cu(L-Phe-Ala)	100	200	-	-	[26]
[Cu(bipy)]Cl <sub>2</sub>	-	60 %*	-	-	[27]
[Cu(dmb)]Cl <sub>2</sub>	-	60 %*	-	-	[27]
[Cu(phen)]Cl <sub>2</sub>	-	10 %*	-	-	[27]
[Cu(4met-phen)]Cl <sub>2</sub>	-	10 %*	-	-	[27]
[Cu(neo)]Cl <sub>2</sub>	-	10 %*	-	-	[27]
[Cu(5-NO <sub>2</sub> -phen) <sub>2</sub> ]Cl <sub>2</sub>	-	5.0	16	-	[28]
[Cu(gly-val)(phen)]	1.0	-	15	14	[29]
[Cu(ala-gly)(phen)]	1.0	-	7.5	9.5	[29]
[Cu(ala-phe)(phen)]	0.94	-	2.0	1.0	[29]
[Cu(phe-ala)(phen)]	13	-	7.0	9.90	[29]
[Cu(phe-val)(phen)]	7.4	-	3.1	7.1	[29]
[Cu(phe-phe)(phen)]	9.6	-	5.2	7.8	[29]
[Cu(Ala-Phe)(5-NO <sub>2</sub> -phen)]	-	4.0	13	-	[28]
[Cu(Phe-Ala)(5-NO <sub>2</sub> -phen)]	-	8.4	>20	-	[28]
[Cu(Phe-Val)(5-NO <sub>2</sub> -phen)]	-	4.8	14	-	[28]
[Cu(Phe-Phe)(5-NO <sub>2</sub> -phen)]	-	9.3	>20	-	[28]
[Cu(ida)(bam)]	-	80 %*	-	-	[27]
[Cu(ida)(bipy)]	-	70 %*	-	-	[27]
[Cu(ida)(dmb)]	-	60 %*	-	-	[27]
[Cu(ida)(phen)]	-	50 %*	-	-	[27]
[Cu(ida)(4met-phen)]	-	30 %*	-	-	[27]
[Cu(ida)(neo)]	-	10 %*	-	-	[27]
[Cu(ida)(batho)]	-	10 %*	-	-	[27]
[CuCl(dmbpy)(PPh <sub>3</sub> )]	-	8.4	4	4.5	[30]
[CuCl(phen)(PPh <sub>3</sub> )]·0.25H <sub>2</sub> O	-	4.3	3.6	3.5	[30]
[CuCl(neo)(PPh <sub>3</sub> )]	-	1.4	2.8	1.3	[30]
Cisplatin	50	50	30	50	[28, 30]

\*Cytotoxic activity expressed as % of cellular viability at 5 mM of the complexes

complexes coordinate to the DNA nucleobases, i.e. they form a covalent bond to the N of the bases, a mechanism similar to that of the Cisplatin. [Cu(L-Ala-Phe)] presented a k<sub>b</sub> of 1x 10<sup>4</sup> [31].

Their cytotoxic activity is lower to that of the Cisplatin [26].

Homoleptic [Cu(diimine)]Cl<sub>2</sub> present varied cytotoxic activity, higher than that of the Cisplatin.

#### [Cu(L-dipeptide)(phen)] and [Cu(L-dipeptide)(5NO<sub>2</sub>-phen)]

Trying to improve the activity of Cu-L-dipeptide complexes, we introduced diimines as second ligands on the CuL-dipeptide complexes, preparing heteroleptic complexes (ie containing two different ligands).

Firstly, Cu-L-dipeptide-phen heteroleptic complexes were prepared. A general scheme of their molecular structure is presented in Figure 2.

The introduction of phen increased the lipophilicity of the complexes, which could improve cellular uptake. Phen works as an ancillary ligand making these complexes intercalate to DNA, as determined by UV-vis, EPR and DC methods.

Cu-L-dipeptide-phen presented high cytotoxic activity, being all more active than Cisplatin (presenting a lower IC50) [29].

Heteroleptic Cu(II) complexes with L-dipeptide and 5-NO<sub>2</sub>-phen were studied to evaluate the influence of the 5-NO<sub>2</sub>-phen moiety in their DNA binding and cytotoxic activity.

DNA binding studies shows that 5-NO<sub>2</sub>-phen (and its complexes) binds to isolated DNA [28]. It also suggests an intercalative mode of binding of the 5-NO<sub>2</sub>-phen and its complexes to the DNA, as observed for other Cu-5-NO<sub>2</sub>-phen heteroleptic complexes [32]. The stacking of the complex molecules possibly occurs through the 5-NO<sub>2</sub>-phen moiety, between the base pairs of DNA. These results are similar to those of the phen ligand and its complexes. The extent of the binding is lower, suggesting that the 5-NO<sub>2</sub> group impairs DNA binding, as determined by CD studies. A similar behavior was reported for Ru heteroleptic complexes [33].

All of the complexes present higher cytotoxic activity than Cisplatin. [Cu(dipeptide)(5-NO<sub>2</sub>-phen)] complexes show similar activity to [Cu(5-NO<sub>2</sub>-phen)<sub>2</sub>]Cl<sub>2</sub>. [Cu(Ala-Phe)(5-NO<sub>2</sub>-phen)] is the more active. Compared with their phen analogs, the 5-NO<sub>2</sub>-phen complexes present lower activity. This tendency was previously observed in the Casiopeinas series, where complexes containing 5-NO<sub>2</sub>-phen had larger IC50 than their counterparts with phen [34], showing the relevance of the phen moiety on the activity of the complexes.

This reduction of the cytotoxic activity is possibly related to the fact that the nitro group impairs DNA binding. The electron-tractor effect of the nitro substituent on the phen can also influence the reactivity of the Cu center in the complexes and therefore their activity.

#### [Cu(ida)(diimine)] complexes

A simpler ligand than dipeptides was assayed as coligand to neutralize the charge of the Cu(diimine)<sup>2+</sup> core. Heteroleptic Cu(iminodiacetate) (diimine) complexes were prepared with a variety of diimines to study their influence in the DNA binding and cytotoxic activity. A general scheme of their molecular structure is presented in Figure 3.

DNA binding results show that all the complexes bind to the DNA, but induce different changes as studied by CD. Complexes [Cu(ida)(bam)(H<sub>2</sub>O)], [Cu(ida)(bipy)] and [Cu(ida)(neo)] bind by partial intercalation to the DNA. Complex [Cu(ida)(5NO<sub>2</sub>-phen)(H<sub>2</sub>O)]

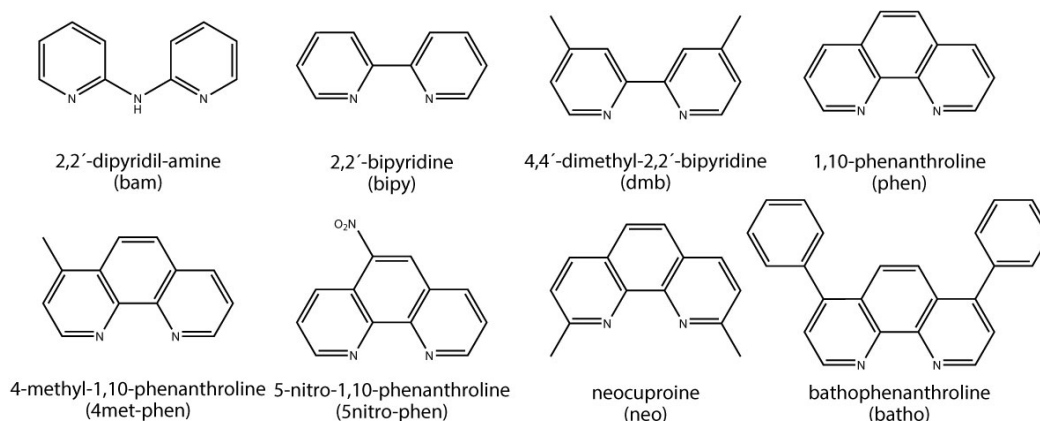


Figure 1. Schematic representation and abbreviation of the different diimines used in this work

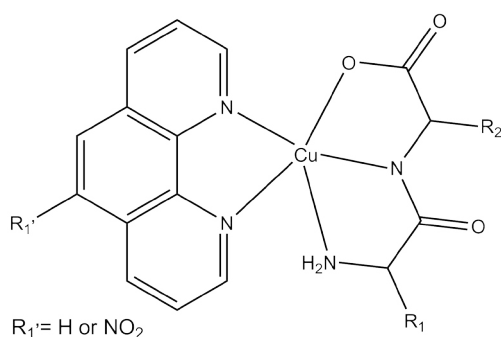


Figure 2. Schematic representation of the molecular structure of the [Cu(L-dipeptide) (phen)] and [Cu(L-dipeptide)(5-NO<sub>2</sub>-phen)] complexes

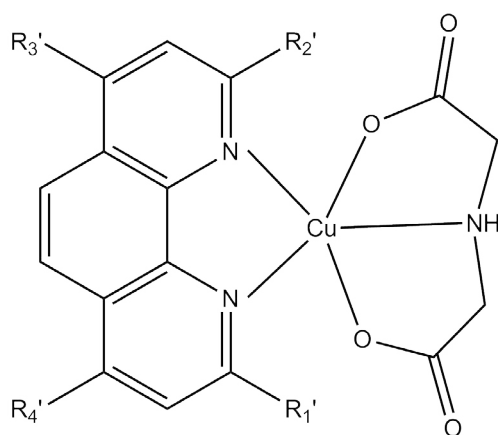


Figure 3. Schematic representation of the molecular structure of the [Cu(ida)(diimine) complexes

induce B to C conformational change on DNA. Complexes [Cu(ida) (phen)(H<sub>2</sub>O)] and [Cu(ida)(4methyl-phen)] induce B to A form conformational change on DNA. Complex [Cu(ida)(batho)] induces conformational change from the right-handed B to the left-handed Z form of the DNA. Their K<sub>b</sub> are in the range 1-3 x10<sup>3</sup>.

As complexes with IC<sub>50</sub> in the lower micromolar range has been considered to show significant cytotoxic activity [7], for these complexes the activity was determined at a fixed dose of 5 μM .

All complexes show increased cytotoxicity than Cisplatin, as observed in Table 1. Among the studied complexes [Cu(ida)(neo)]

and [Cu(ida)(batho)] present the higher activity, comparable to those presented by other complexes catalogued as promising antitumor agents [25,35-37]. No correlation can be stated between the cytotoxic activity and DNA binding (K<sub>b</sub> or induced conformational changes). Cu(ida)(4met-phen), [Cu(ida)(phen)] and [Cu(ida)(batho)] complexes present significant cytotoxic activity. The order of activity is in agreement with results observed in other series of Cu-diimine mixed ligand complexes [34].

In other series of Cu(II) mixed ligand complexes containing diimines, a similar relation between the aromatic di-imine ligand and the cytotoxic activity was observed [34].

The cytotoxic activity of the heteroleptic complexes is similar to that of the corresponding homoleptic [Cu(diimine)Cl<sub>2</sub>] complex. It suggests that *in vivo* the complexes are dissociated to the respective [Cu(diimine)] complex. This is in agreement with the expected reduction of Cu (II) to Cu(I) inside the cell, which favors the release of the ida ligand. Therefore, the active species could be the cationic ones.

### Copper (I) complexes of diimines

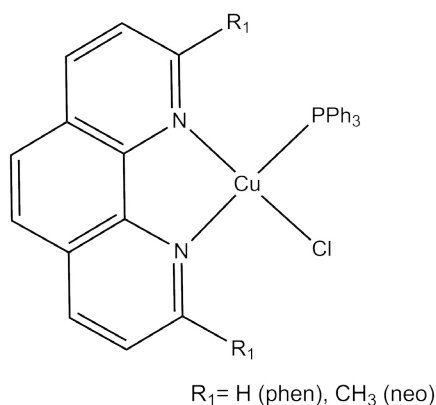
Intracellular copper is in the chemical form of Cu(I); Cu(II) complexes are believed to be reduced to Cu(I) in the intracellular media [7,24,38]. Despite that, few studies report on the antitumor activity of Cu(I) complexes, probably due to the difficulty in obtaining stable Cu(I) complexes in aqueous media. Encouraged by the activity and low toxicity of HydroCuP<sup>+</sup> we explored the cytotoxic activity of Cu-diimine systems adding phosphine ligands to stabilize Cu(I). A general scheme of their molecular structure is presented in Figure 4.

The complexes degrade DNA rapidly therefore no binding constant could be determined.

All the complexes present higher cytotoxic activity than Cisplatin and according to Santini *et al.* [7] may be classified as very active (IC<sub>50</sub> in the low micromolar range). The cytotoxic activity increases in the same order as observed for the lipophilicity, being [Cu(neo)(PPh<sub>3</sub>)Cl] the most lipophilic and active. This observation is in agreement with the hypothesis that the complexes act as "copper ionophores" that deliver copper inside the cell [39,40].

### Proposed mechanism of action of the complexes

As mentioned in Section 1, the mechanisms underlying the biological activity of Cu-complexes are not completely known [14,15].



**Figure 4.** Schematic representation of the molecular structure of the [Cu(diimine)(phosfine)Cl] complexes

Taken our results into account and the literature it can be proposed that the molecular events that cause the cytotoxic activity of Cu-diimine complexes include:

- Intercalation to the DNA inducing different conformational changes or even DNA degradation. Only Cu-L-dipeptide complexes coordinate to DNA as Cisplatin.
- Generation of ROS which oxidize and degrade biomolecules, as DNA [41]. ROS production *in vitro* has been observed for related complexes.
- Mitochondrial toxicity as observed in the literature for related Cu-complexes [42].

In the case of the [Cu(ida)(batho)] complex, the more lipophilic one, it is possible that the activity is due to the formation of a lipophilic delocalized cation [Cu(batho)]<sup>2+</sup>, which are known to selectively accumulate in mitochondria of cancer cells [43,44]. These effect can't be ruled out for the rest of the compounds, that is, the active species for these complexes can be the cation (more or less lipofilic) [Cu(diimine)]<sup>2+</sup>.

With relation to the cellular death mechanism, for related copper complexes cellular death by an apoptotic mechanism was observed [7,42]. Despite that, cellular death by a paraptotic mechanism can't be ruled out as it was reported for Cu-phosphine complexes [12,39].

## Conclusions and perspectives

All complexes bind to DNA, being the diimine the ligand which most influenced the DNA binding. All the complexes were active against the tested tumor cell lines, most of them with increased activity than the reference drug Cisplatin. The diimine plays a central role in the anticancer activity, possible being the [Cu(diimine)]<sup>+</sup> the active species inside the cell. The anionic ligand modulates the activity. There is no apparent relationship between DNA binding and cytotoxic activity.

Among all the dipeptide containing complexes, those containing L-Ala-Phe are the more active, although the dipeptide itself doesn't present cytotoxic activity. This fact is probably related to the major stability of the complex. [Cu(ala-phe)(phen)], [CuCl(neo)(PPh<sub>3</sub>)], [Cu(ida)(neo)], [Cu(ida)(batho)] and [Cu(ida)(4met-phen)] showed the strongest cytotoxic activity against cancer cell lines and therefore, those compounds are good candidates to test its antitumor activity *in vivo*.

More research is needed in order to establish structure – activity relationship and advance towards the rational design of Copper-diimine complexes with antitumor activity.

*In vivo* studies are necessary to advance in the potential use of these compounds as metallo drugs, especially to gain insight in their selectivity and eventual cure rates which are not accessible by *in vitro* assays.

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